**Karan Arora**  **R.L. Institute M: 9416974837**

**Max Time : 1 hr** **Class = 10th Science Test Max Marks : 20**

**METAL and NON-METALS**

1. Multiple choice questions : [ 1 X 4 = 4 ]
2. Galvanisation is a method of protecting iron from rusting by coating with a thin layer of :

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| a) Gallium | b) Zinc | c) Silver | d) Aluminum |

1. If copper is kept open in air, it slowly loses its shining brown surface and gains a green coating. It is due to formation of :

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| a) Cu(NO3)2 | b) CuCO3 | c) CuSO4 | d) CuO |

1. The metal which is known as strategic metal is :

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| a) Manganese | b) Zirconium | c) Uranium | d) all of these |

1. Which among the following alloys containing mercury as one of its components?

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| a) Alnico | b) Solder | c) Stainless steel | d) Zinc amalgam |

1. Explain Anodising. [ 1 ]
2. Write the chemical formula of : (i) Cinnabar (ii) Calamine. [ 1 ]
3. Name two metals, which are found in nature in the free state. [ 1 ]
4. Why is stainless steel preferred for making household utensils? [ 1 ]
5. What is Galvanisation. [ 1 ]
6. Name the alloy used for welding electric wires together and write its constituents. [ 1 ]
7. Give the reaction involved during extraction of zinc from its ore by : [ 2 ]

(i) Roasting of zinc ore (ii) Calcination of zinc ore

1. (a) What is thermit process? Where is this process used? Write balanced chemical equation

involved in the reaction. [ 2 ]

(b) Where does the metal aluminium, used in the process, occurs in the reactivity series of

metals?

(c) Name the substances that are getting oxidized and reduced in the process.

1. (a) Distinguish between roasting and calcination. [ 3 ]

(b) Write a balanced chemical equation to illustrate the use of aluminium for joining cracked

railway lines.

1. Explain corrosion with example. Write 3 methods of prevention of corrosion of iron. [ 3 ]

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